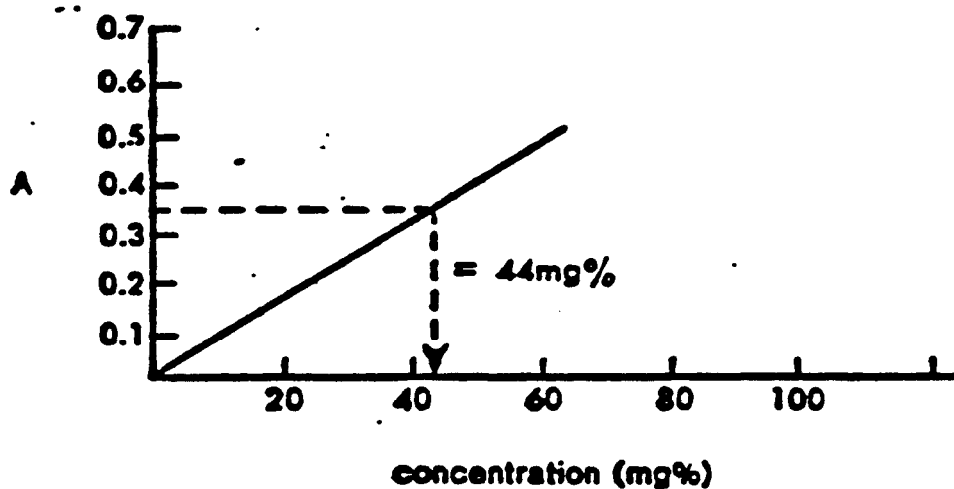


This Standard Curve can now be used to convert the light absorbances of test solutions into their respective concentrations. As an example, if a solution of unknown concentration has an absorbance of 0.35, use the graph to determine what its concentration would be.



Again, the advantage of using a Standard Curve rather than a single known (standard) solution whose value is "plugged into" the Beer's Law Equation is that it "averages-out" experimental error.